The Pressure of A + Pressure of B = 3 Bar (Given :Pressure when both gases are present in the flask under same conditions is 3bar)

Using Daltons law of partial Pressures,
Pressure of B = 3 – 2 = 1 bar
According to an ideal gas equation,
For gas A:

$$P_1V=n_1RT \implies P_1V = RT\frac{m1}{M}$$

Similarly for the gas B:

$$P_2V=n_2RT \implies P_2V = RT\frac{m2}{M2}$$

Diving both the Equ. We get:

$$\frac{P*P1}{P2} = \frac{m\,1\,M2}{m\,2\,M\,1} \implies \frac{P1m2}{P2m1} = \frac{M2}{M1} = \frac{P1m2}{P2m1}$$

So the ratio of their

Molecular masses:

$$\frac{M2}{M1} = \frac{2*2}{1*1} = 4{:}1$$