

The constancy of the solubility product is derived from the law of mass action. It represents a specific application of this law to the case of the equilibrium between an electrolyte in the solid phase and a saturated solution of the same electrolyte. It is assumed that the electrolyte in the solution is fully dissociated. The solubility product is most accurately measured by the electromotive force (emf) method; another common method of measurement derives from the determination of solubility by the electrical conductivity of saturated solutions. The solubility product for many compounds is established with sufficient accuracy for all practical purposes. In tables, the solubility product is usually given for a temperature of 25°C (sometimes for 18°C).