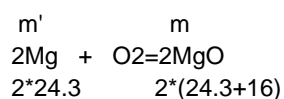


To calculate the mass or percent error you need to do the following:

First: Find procedural error that affects the mass of magnesium oxide:  
So if Mg reacted incompletely the mass of MgO is lower than theoretical:

W is "yeild"  
m' is mass of Mg  
m is mass of MgO  
m" is mass of O

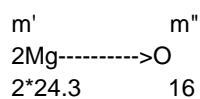


$$m = 40.3m' / 24.3 = 1.66m'$$

If reacted 90% you'll get  $m = 0.9 \cdot 1.66m'$

Second: Find procedural error that affects the mass of oxygen in magnesium oxide:  
Now you need to conect unreacted Mg with O from MgO

$2\text{Mg} + \text{O}_2 = 2\text{MgO}$   
The scheme of this process is:



$$m'' = 16 \cdot m' / 2 \cdot 24.3 = 0.33m'$$