How many atoms are there in 71.9 g of nickel.

Solbtion:

We use the Avogadro's number: $6.02 \cdot 10^{23}$ atoms are included to 1 mole of some compound.

The mass of 1 mole of Ni is 58.69 g, so 71.9 g of Ni include $\frac{6.02 \cdot 10^{23} \cdot 71.9}{58.69} = 7.37 \cdot 10^{23}\,\text{atoms}$

Answer:

The number of atoms in 71.9 g of nickel are $7.37 \cdot 10^{23}$.