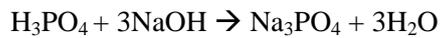How many mL of 0.75 mol/L solution of phosphoric acid will be neutralized by 25 mL of 0.5 mol/L solution of sodium hydroxide?

$H_3PO_4 + 3NaOH \rightarrow Na_3PO_4 + 3H_2O$

ν (NaOH) = V*c;
ν (NaOH) = 0.025 * 0.5 = 0.0125 mol
ν (H_3PO_4) = 0.0125 / 3 = 0.0042 mol
V (H_3PO_4) = ν / c;
V (H_3PO_4) = 0.0042 / 0.75 = 0.0055 L

V (H_3PO_4) = 0.0055 L