convert to moles: 458166 atoms of H20

if 458166 is number of molecules H2O $v = 7.6*10^{-19}mol$

$$v = N\big/N_A$$
$$v = \frac{458166}{602000*10^{18}} = 7.6*10^{-19}mol$$

If 458166 is number of H+H+O atoms → number of molecules = 458166/3 = 152722→
$v = 2.5*10^{-19}mol$

$$v = N\big/N_A$$
$$v = \frac{152722}{602000*10^{18}} = 2.5*10^{-19}mol$$