

2.61 g of  $\text{KClO}_3$  were present in a 5.50 g sample of  $\text{KClO}_3$  -  $\text{KCl}$  mixture, what is the percent by mass of  $\text{KClO}_3$  in the mixture?

also please show percent if the 2.61 were 0.904

$$\omega = \frac{m(\text{KClO}_3)}{m(\text{KClO}_3) + m(\text{KCl})} * 100\%$$

$$\omega = \frac{2.61}{5.50} * 100\% = 47.45\% \text{ wt}$$

$$\underline{W(\text{KClO}_3) = 47.45\% \text{ wt}}$$

if 2.61 of  $\text{KClO}_3 = 90.4\%$

$$W = 100\% - 90.4\% = 9.6\%$$

$$\underline{W = 9.6\%}$$