In accordance with Ideal gas law we have the following equation:

nRT = pV, where

P is the pressure of the gas, V is the volume of the gas, n is the amount of substance of gas (also known as number of moles), T is the temperature of the gas and R is the ideal, or universal, gas constant.

$R = 8.314\ J \cdot K^{-1} \cdot mol^{-1}$.

First step will be the following:

$n*8.314\ J \cdot K^{-1} \cdot mol^{-1}*(100+273)K = 775*133,322Pa*6,75*10^{-3}m^3$

Hence, n = 0,225 moles.

We use the Ideal gas law for determination the gas volume at STP:

$0,225mol*8.314\ J \cdot K^{-1} \cdot mol^{-1}*273K = 760*133,322Pa*V$

From this equation we have, that gas volume is $0,005m^3$ or 5 liters.